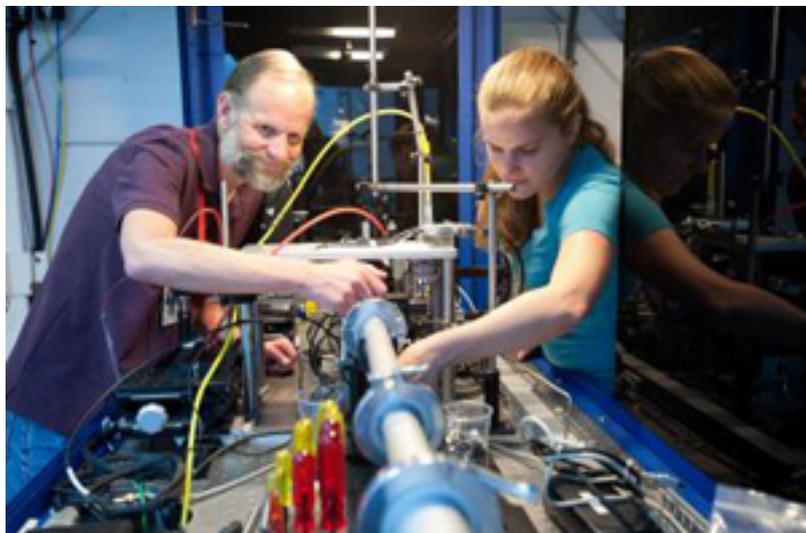


Stanford-SLAC team uses X-ray imaging to observe running batteries in action



Mike Toney and Johanna Nelson demonstrate the high-power transmission X-ray Microscope at SLAC's Stanford Synchrotron Radiation Lightsource.

Most electric cars, from the Tesla Model S to the Nissan Leaf, run on rechargeable lithium-ion batteries – a pricey technology that accounts for more than half of the vehicle's total cost. One promising alternative is the lithium-sulfur battery, which can theoretically store five times more energy at a much lower cost.

But lithium-sulfur technology has a major drawback: After a few dozen cycles of charging and discharging, the battery stops working.

"The cycle life of lithium-sulfur batteries is very short," said Johanna Nelson, a postdoctoral scholar at the SLAC National Accelerator Laboratory at Stanford University. "Typically, after a few tens of cycles the battery will die, so it isn't viable for electric vehicles, which require many thousands of cycles over a 10- or 20-year lifetime."

A typical lithium-sulfur battery consists of two electrodes – a lithium metal anode and a sulfur-carbon cathode – surrounded by a conductive fluid, or electrolyte. Several studies have attributed the battery's short cycle life to chemical reactions that deplete the cathode of sulfur.

But a recent study by Nelson and her colleagues is raising doubts about the validity of previous experiments. Using high-power X-ray imaging of an actual working battery, the Stanford-SLAC team discovered that sulfur particles in the cathode largely remain intact during discharge. Their results, published in the *Journal of the American Chemical Society* (JACS), could help scientists find new ways to develop commercially viable lithium-sulfur batteries for electric vehicles.

"Based on previous experiments, we expected sulfur particles to completely disappear from the cathode when the battery discharges," said Nelson, the lead author of the JACS study. "Instead, we saw only negligible changes in the size of the particles, the exact opposite of what earlier studies found."

Nelson and her co-workers conducted their experiments at SLAC using two powerful imaging techniques: X-ray diffraction and transmission X-ray microscopy. The X-ray microscope enabled the researchers to take nanosize snapshots of individual sulfur particles before, during and after discharge – the first real-time imaging of a lithium-sulfur battery in operation.

"The standard way to do high-resolution imaging is with electron microscopes after the battery has partially discharged," Nelson said. "But electrons don't penetrate metal and plastic very well. With SLAC's X-ray microscope, we can actually see changes that are happening while the battery is running."

Pesky polysulfides

In lithium-sulfur batteries, an electric current is generated when lithium ions in the anode react with sulfur particles at the cathode during discharge. The byproducts of this chemical reaction are compounds known as lithium polysulfides.

Problems can arise when the polysulfides leak into the electrolyte and permanently bond with the lithium metal anode. "When that happens, all of the sulfur material in the polysulfides is lost," Nelson said. "It will never recycle. You don't want to lose active sulfur material every time the battery discharges. You want a battery that can be cycled multiple times."

Previous experiments also showed the formation of dilithium sulfide (Li_2S) crystals during the discharge phase. "Crystalline Li_2S and polysulfides can form a thin film that prevents the conduction of electrons and lithium ions," Nelson said. "The film acts as an insulating layer that can cause the battery to die."

Several studies using electron microscopes produced images of electrodes coated with polysulfides and crystalline Li_2S , and cathodes depleted of sulfur. Those images led researchers to conclude that much of the sulfur had been chemically transformed into Li_2S -polysulfide sheets that prevented the battery from operating.

Flawed findings

But according to Nelson and her colleagues, some of the previous studies were flawed. "The approach they were using was mistaken," Nelson said. "Typically, they would cycle the battery, disassemble it, wash away the electrolyte and then analyze it with X-ray diffraction or an electron microscope. But when you do that, you also wash away all of the polysulfides that are loosely trapped on the cathode. So when you image the cathode, you don't see any sulfur species at all."

The Stanford-SLAC team took a different approach. Researchers used the

transmission X-ray microscope at SLAC to take multiple images of tiny sulfur particles every five minutes while the battery discharged. Each particle was a fraction of the size of a grain of sand. The results were clear: Every particle retained its basic shape and size throughout the discharge cycle.

"We expected the sulfur to completely disappear and form polysulfides in the electrolyte," Nelson said. "Instead we found that, for the most part, the particles stayed where they were and lost very little mass. They did form polysulfides, but most of those were trapped near the carbon-sulfur cathode. We didn't have to disassemble the battery or even stop it, because we could image the sulfur content while the device was operating."

X-ray diffraction yielded an additional surprise. "Based on previous experiments, we expected that crystalline Li_2S would form at the end of the discharge cycle," she said. "But we did a very deep discharge and never saw any Li_2S in its crystalline state."

Future research

The Stanford-SLAC study could open new avenues of research that could improve the performance of lithium-sulfur batteries, said co-author Michael Toney, head of the Materials Sciences Division at SLAC's Stanford Synchrotron Radiation Lightsource.

"Our study demonstrates the importance of using high-power X-ray technologies to study batteries while they are operating," Toney said. "From an engineering standpoint, it's valuable to know that relying on standard electron microscopy to test the fidelity of materials may give you deceptive results."

Several research labs are looking for new ways to trap polysulfides on the cathode. A variety of techniques have shown promise, including novel electrolytes and carbon nanotubes coated with sulfur.

But the polysulfide problem might not be as daunting as previous studies suggest.

"We found that very few of the polysulfides went into the electrolyte," Nelson said. "The carbon-sulfur cathode actually trapped them better than expected. But even a small amount of polysulfides will cause the battery to fail within 10 cycles. If scientists want to improve the cycle life of the battery, they need to prevent virtually all of the polysulfides from leaking into the electrolyte. If they really want to know what's going on inside the battery, they can't just use standard analysis. They need a technology that tells the whole story."

###

In addition to Nelson, the co-lead authors of the JACS study are SLAC postdoctoral researcher Sumohan Misra and Stanford doctoral student Yuan Yang.

The study is also co-authored by Yi Cui, an associate professor of materials science and engineering at Stanford and of photon science at SLAC; Hongjie Dai, a professor

Stanford-SLAC team uses X-ray imaging to observe running batteries in act

Published on Electronic Component News (<http://www.ecnmag.com>)

of chemistry at Stanford; graduate students Ariel Jackson and Hailiang Wang of Stanford; and Joy C. Andrews, a staff scientist at SLAC.

The research was supported by the Department of Energy, the Department of Defense and a Stanford Graduate Fellowship.

SLAC is a national laboratory operated by Stanford for the DOE. The study was conducted in cooperation with the Stanford Institute for Materials and Energy Science, a Stanford-SLAC research partnership.

This article was written by Mark Schwartz, Precourt Institute for Energy at Stanford University.

Related information:

[Toney Research Lab](#) [1]

[Stanford Synchrotron Radiation Lightsource](#) [2]

[Stanford Institute for Materials and Energy Science](#) [3]

Original release: http://www.eurekalert.org/pub_releases/2012-07/su-stu071812.php [4]

Source URL (retrieved on 07/29/2014 - 5:42pm):

<http://www.ecnmag.com/news/2012/07/stanford-slac-team-uses-x-ray-imaging-observe-running-batteries-action>

Links:

[1] <http://www-ssrl.slac.stanford.edu/toneygroup/index.html>

[2] <http://www-ssrl.slac.stanford.edu/>

[3] <http://simes.slac.stanford.edu/>

[4] http://www.eurekalert.org/pub_releases/2012-07/su-stu071812.php